

3学年理系

家庭學習期間

化学課題解答

【採点して登校日に課題提出】

- (1) $Zn + H_2SO_4 \rightarrow ZnSO_4 + H_2 \uparrow$
- (2) $CaF_2 + H_2SO_4 \xrightarrow{\text{加熱}} CaSO_4 + 2HF \uparrow$
- (3) $2F_2 + H_2O \rightarrow 4HF + O_2 \uparrow$
- (4) $SiO_2 + 6HF \rightarrow H_2SiF_6 + 2H_2O$
- (5) $4HCl + MnO_2 \xrightarrow{\text{加熱}} MnCl_2 + 2H_2O + Cl_2 \uparrow$
- (6) $CaCl(ClO) \cdot H_2O + 2HCl \rightarrow CaCl_2 + 2H_2O + Cl_2 \uparrow$
- (7) $Cl_2 + H_2O \rightleftharpoons HCl + HClO$
- (8) $Cl_2 + Ca(OH)_2 \rightarrow CaCl(ClO) \cdot H_2O$
- (9) $Cl_2 + H_2 \xrightarrow{\text{光}} 2HCl$
- (10) $NaCl + H_2SO_4 \xrightarrow{\text{加熱}} NaHSO_4 + HCl \uparrow$

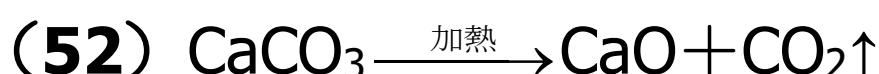
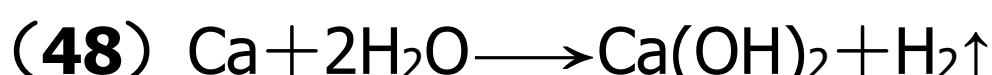
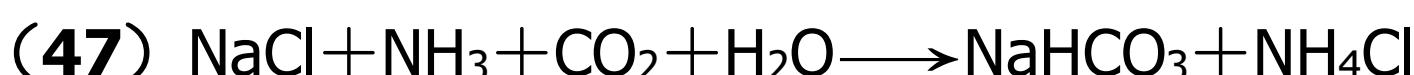
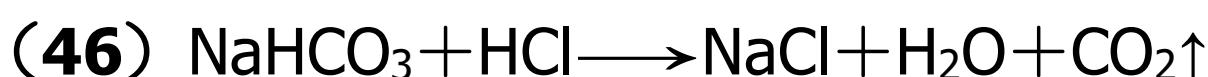
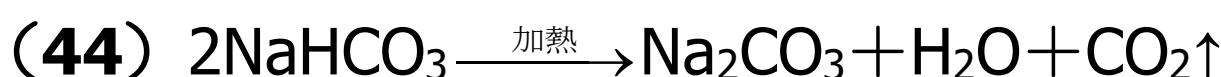
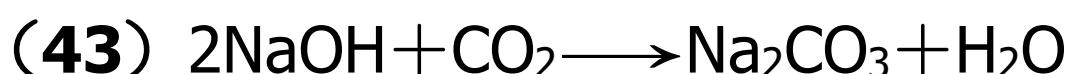
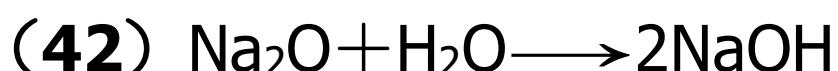
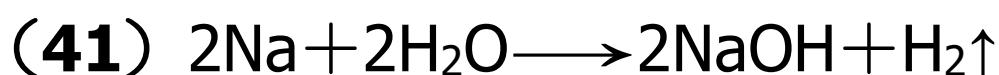
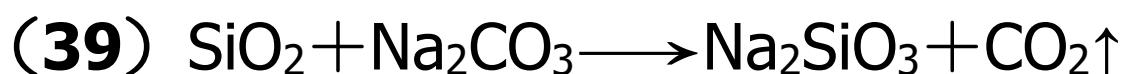
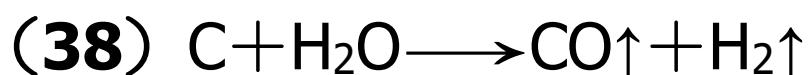
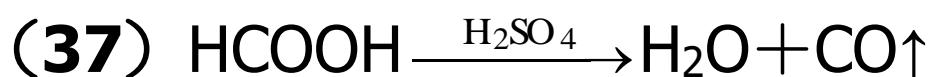
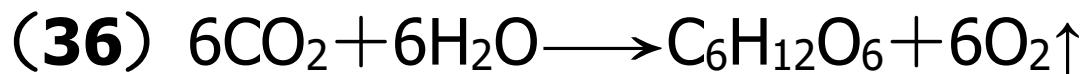
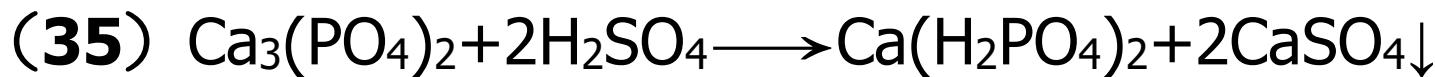
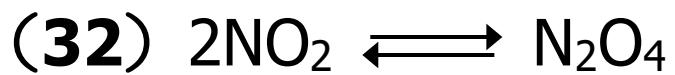
化学反応式

解答欄

- (11) $2\text{KBr} + \text{Cl}_2 \rightarrow 2\text{KCl} + \text{Br}_2$
- (12) $2\text{KI} + \text{Cl}_2 \rightarrow 2\text{KCl} + \text{I}_2 \downarrow$
- (13) $2\text{KI} + \text{Br}_2 \rightarrow 2\text{KBr} + \text{I}_2 \downarrow$
- (14) $2\text{KClO}_3 \xrightarrow{\text{MnO}_2} 2\text{KCl} + 3\text{O}_2 \uparrow$
- (15) $2\text{H}_2\text{O}_2 \xrightarrow{\text{MnO}_2} 2\text{H}_2\text{O} + \text{O}_2 \uparrow$
- (16) $2\text{KI} + \text{H}_2\text{O} + \text{O}_3 \rightarrow 2\text{KOH} + \text{O}_2 \uparrow + \text{I}_2 \downarrow$
- (17) $\text{S} + \text{O}_2 \rightarrow \text{SO}_2 \uparrow$
- (18) $2\text{SO}_2 + \text{O}_2 \xrightarrow{\text{V}_2\text{O}_5} 2\text{SO}_3$
- (19) $\text{SO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_4$
- (20) $\text{NaHSO}_3 + \text{H}_2\text{SO}_4 \rightarrow \text{NaHSO}_4 + \text{H}_2\text{O} + \text{SO}_2 \uparrow$
- (21) $\text{Cu} + 2\text{H}_2\text{SO}_4 \xrightarrow{\text{加熱}} \text{CuSO}_4 + 2\text{H}_2\text{O} + \text{SO}_2 \uparrow$
- (22) $\text{FeS} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2\text{S} \uparrow$
- (23) $\text{Cu} + 4\text{HNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{H}_2\text{O} + 2\text{NO}_2 \uparrow$
- (24) $3\text{Cu} + 8\text{HNO}_3 \rightarrow 3\text{Cu}(\text{NO}_3)_2 + 4\text{H}_2\text{O} + 2\text{NO} \uparrow$
- (25) $2\text{NH}_4\text{Cl} + \text{Ca}(\text{OH})_2 \xrightarrow{\text{加熱}} \text{CaCl}_2 + 2\text{H}_2\text{O} + 2\text{NH}_3$
- (26) $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$
- (27) $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$
- (28) $\text{NH}_3 + \text{H}_2\text{O} \rightleftharpoons \text{NH}_4^+ + \text{OH}^-$
- (29) $4\text{NH}_3 + 5\text{O}_2 \xrightarrow{\text{Pt}} 4\text{NO} \uparrow + 6\text{H}_2\text{O}$
- (30) $2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2 \uparrow$
- (31) $3\text{NO}_2 + \text{H}_2\text{O} \rightarrow 2\text{HNO}_3 + \text{NO} \uparrow$

化学反応式

解答欄



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- (53) $\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2 \uparrow$
- (54) $\text{CaC}_2 + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{C}_2\text{H}_2 \uparrow$
- (55) $2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2 \uparrow$
- (56) $2\text{Al} + 2\text{NaOH} + 6\text{H}_2\text{O} \rightarrow 2\text{Na}[\text{Al(OH)}_4] + 3\text{H}_2 \uparrow$
- (57) $\text{Al}_2\text{O}_3 + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2\text{O}$
- (58) $\text{Al}_2\text{O}_3 + 2\text{NaOH} + 3\text{H}_2\text{O} \rightarrow 2\text{Na}[\text{Al(OH)}_4]$
- (59) $\text{Al(OH)}_3 + 3\text{HCl} \rightarrow \text{AlCl}_3 + 3\text{H}_2\text{O}$
- (60) $\text{Al(OH)}_3 + \text{NaOH} \rightarrow \text{Na}[\text{Al(OH)}_4]$
- (61) $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2 \uparrow$
- (62) $\text{Zn} + 2\text{NaOH} + 2\text{H}_2\text{O} \rightarrow \text{Na}_2[\text{Zn(OH)}_4] + \text{H}_2 \uparrow$
- (63) $\text{ZnO} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2\text{O}$
- (64) $\text{ZnO} + 2\text{NaOH} + \text{H}_2\text{O} \rightarrow \text{Na}_2[\text{Zn(OH)}_4]$
- (65) $\text{Zn(OH)}_2 + 2\text{HCl} \rightarrow \text{ZnCl}_2 + 2\text{H}_2\text{O}$
- (66) $\text{Zn(OH)}_2 + 2\text{NaOH} \rightarrow \text{Na}_2[\text{Zn(OH)}_4]$
- (67) $\text{Cu(OH)}_2 + 4\text{NH}_3 \rightarrow [\text{Cu}(\text{NH}_3)_4]^{2+} + 2\text{OH}^-$
- (68) $\text{CuSO}_4 + \text{H}_2\text{S} \rightarrow \text{CuS} \downarrow + \text{H}_2\text{SO}_4$
- (69) $2\text{Ag}^+ + 2\text{OH}^- \rightarrow \text{Ag}_2\text{O} \downarrow + \text{H}_2\text{O}$
- (70) $\text{Ag}_2\text{O} + 4\text{NH}_3 + \text{H}_2\text{O} \rightarrow 2[\text{Ag}(\text{NH}_3)_2]^+ + 2\text{OH}^-$
- (71) $\text{Fe}_2\text{O}_3 + 2\text{Al} \rightarrow \text{Al}_2\text{O}_3 + 2\text{Fe}$
- (72) $\text{FeCl}_3 + 3\text{H}_2\text{O} \rightarrow \text{Fe(OH)}_3 + 3\text{HCl}$
- (73) $\text{Pb} + \text{PbO}_2 + 2\text{H}_2\text{SO}_4 \rightleftharpoons 2\text{PbSO}_4 \downarrow + 2\text{H}_2\text{O}$